

## PROPELLANT FOR GAS GENERATORS

### BACKGROUND OF THE INVENTION

Gas generators have become of increasing interest for lifesaving purposes, for example in vehicles. Worldwide, the most popular mixture for gas generation contains sodium azide. However, sodium azide is toxic, requiring special measures during the manufacture of the raw material, during the preparation of the gas charge composition, and during its processing, quality control, and waste removal. This holds true, in particular, in connection with the scrapping of vehicles.

There has been a large number of attempts to utilize substances other than sodium azide. Thus, DE-A-2,142,578 describes a press-molded propellant charge for the rapid inflation of a hollow body by the reaction of tetrazylazene with oxygen carriers. DE-A-1,806,550 proposes a propellant charge which generates pressure gas, yielding cool gases, based on ammonium nitrate, activated carbon, and a compound that is subject to endothermal decomposition or sublimation. However, this system produces a large proportion of steam, representing a drawback since water leads to a strong increase in temperature due to its high heat of condensation.

DE-A-1,222,418 discloses mixtures generating pressurized gas, based on inorganic perchlorate oxidizers, polymeric fuel binders, and a coolant. Preparations having high proportions of chlorate or perchlorate, however, lead to chlorine components in

the reaction gases. Thus, EP-A-372,733 likewise yields an unsatisfactory mixture inasmuch as the propellant charge for the proposed airbag contains about 40% ammonium perchlorate. Even nitrocellulose and nitroglycerin compositions can be found in the literature. Such suggestions cannot be considered for use in lifesaving systems. Nitrocellulose and nitroglycerin mixtures, or also other energetic compounds rich in carbon must be eliminated on account of carbon monoxide formation.

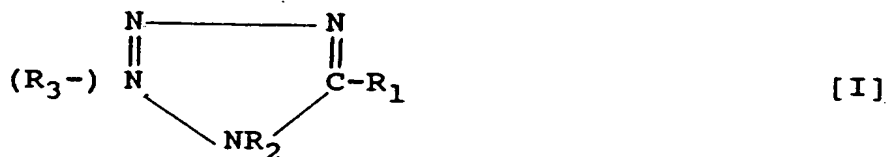
The propellant charges in DE-A-1,250,318, containing aminotetrazole, potassium dichromate, calcium resinate, and metallic silicon, do not meet up-to-date safety requirements, either. The same holds true for DE-C-2,004,620 wherein the compressed gas-producing charges contain azotetrazole and/or ditetrazole and chlorates or perchlorates. The propellant charges of U.S. A-3,734,789 which contain 5-aminotetrazole nitrate and polyisoprene binders, although burning up rapidly, also generate carbon monoxide owing to the carbon-rich binder proportion, in health-endangering concentrations.

#### SUMMARY OF THE INVENTION

Accordingly, the invention is based on the object of making available gas compositions, the manufacture and processing and/or handling of which are harmless and the reaction products of which are nontoxic.

This object has been attained by a propellant or gas generating composition for gas generators containing, as the

nitrogenous compound, (a) tetrazol or one or more tetrazole  
derivatives of the following formula I



wherein  $\text{R}_2$  and  $\text{R}_3$  can be identical or different (but either  
 $\text{R}_2$  or  $\text{R}_3$  is present) and can be: hydrogen, hydroxy, amino,  
carboxy, an alkyl residue of 1-7 carbon atoms, an alkenyl residue  
of 2-7 carbon atoms, an alkylamino residue of 1-10 carbon atoms,  
an aryl residue optionally substituted by one or several  
substituents which can be identical or different, selected from  
the amino group, the nitro group, the alkyl residues of 1-4  
carbon atoms, or an arylamino residue wherein the aryl residue  
can be optionally substituted, or their sodium, potassium, and  
guanidinium salts, or containing, as the nitrogenous compounds,

(b) respectively one or several compounds from the groups of

(A) cyanic acid derivatives and their salts,

(B) triazine and triazine derivatives,

(C) urea, its salts, derivatives, and compounds evolved

therefrom and their salts,

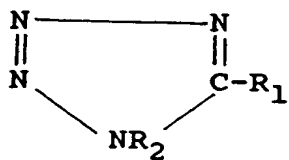
wherein the compounds recited under (a) and (b) can also be  
present as mixtures,

and containing an oxidizing agent from the group of the  
peroxides, or from the group of the peroxides together with

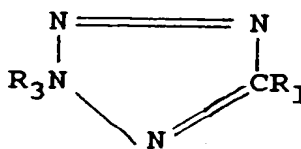
oxidizing agents from the group of the nitrates.

#### DETAILED DESCRIPTION OF THE INVENTION

By the formula I, applicants are referring to tetrazole or derivatives of tetrazole of either the 1-H or 2-H tautomeric forms as shown in the following formulae IA and IB, respectively:



IA



IB

The nitrogen-containing compounds to be used according to this invention are those forming, in a mixture with oxidizing agents, during their thermal-chemical reaction, primarily CO<sub>2</sub>, N<sub>2</sub> and H<sub>2</sub>O, but not releasing any gases, such as CO or NO<sub>x</sub> in health-endangering concentrations. An especially significant feature resides in that the addition of binders is not absolutely necessary.

R<sub>1</sub> preferably is hydrogen, amino, hydroxy, carboxy, a methyl, ethyl, propyl or isopropyl, butyl, isobutyl or tert-butyl, n-pentyl, n-hexyl or n-heptyl residue, a methylamino, ethylamino, dimethylamino, n-heptylamino, n-octylamino or n-decylamino residue, a phenylamino residue, or a phenyl or aminophenyl residue.

$R_2$  or  $R_3$  is preferably hydrogen, a methyl or ethyl residue, or a phenyl or aminophenyl residue.

The following tetrazole derivatives are especially preferred:

5 5-Aminotetrazole, sodium, potassium or calcium

5-aminotetrazolate, 1-(4-aminophenyl)tetrazole,

1-methyl-5-dimethylaminotetrazole, 1-methyl-5-methyl-

aminotetrazole, 1-methyltetrazole, 1-phenyl-5-aminotetrazole,

1-phenyl-5-hydroxytetrazole, 1-phenyltetrazole,

10 2-ethyl-5-aminotetrazole, 2-methyl-5-aminotetrazole,

2-methyl-5-carboxytetrazole, 2-methyl-5-methylaminotetrazole,

2-methyltetrazole, 2-phenyltetrazole, 5-(p-tolyl)tetrazole,

5-diallylaminotetrazole, 5-dimethylaminotetrazole,

5-ethylaminotetrazole, 5-hydroxytetrazole, 5-methyltetrazole,

15 5-methylaminotetrazole, 5-n-decylaminotetrazole,

5-n-heptylaminotetrazole, 5-n-octylaminotetrazole,

5-phenyltetrazole, 5-phenylaminotetrazole, or

bis(aminoguanidine)azotetrazole.

Cyanic acid derivatives used with preference are sodium

20 cyanate, cyanuric acid, 1-cyanoguanidine and/or disodium

cyanamide; triazine derivatives used with preference are

1,3,5-triazine, cyanuric acid esters and/or cyanuric acid amide

(melamine); and urea derivatives used with preference are biuret,

guanidine, nitroguanidine, guanidine nitrate, aminoguanidine,

25 aminoguanidine nitrate, triaminoguanidine nitrate, aminoguanidine

hydrogen carbonate, azodicarboxylic acid diamide, dicyandiamidine

nitrate, dicyandiamidine sulfate, tetrazene and/or semicarbazide nitrate.

5 B side s being acceptable from health viewpoints, the mixtures according to this invention exhibit a high thermal and weather stability, which is a prerequisite for perfect action even after long-term storage.

Suitable oxidizing agents are nitrates of ammonium, sodium, potassium, magnesium, calcium or iron, preferably sodium nitrate, or peroxides of zinc, calcium, strontium or magnesium.

10 The peroxides are used with an oxygen value as obtainable from stable compounds. For zinc peroxide, this value is about 11 to 14% by weight. The corresponding molar ratio of nitrogen-containing compound to peroxide is in a range from 1 : 2 to 5.5. Calcium peroxide can exhibit an active oxygen value of, 15 for example, 18.26% by weight, and grain sizes of 15.5  $\mu\text{m}$  and is advantageously utilized in a molar ratio of nitrogen-containing compound/peroxide of 1 : 3. In general, the aforementioned peroxides can be used in a molar ratio of nitrogen-containing compound to peroxide in a range of 1 : 1 to 20.

20 Calcium peroxide and/or zinc peroxide is used with preference. It is also possible to employ mixtures of the peroxides with one another or mixtures with other oxidizing agents. Other oxidizing agents are, for example, the above-mentioned nitrates of ammonium, sodium, potassium, 25 magnesium, calcium or iron, preferably sodium nitrate.

When using magnesium peroxide and, in particular, calcium or strontium peroxide, the alkaline-acting hydrolysis products may evoke reactions with the other components of the mixture. In this case, coating of the peroxides with inorganic or organic materials in accordance with conventional methods is suitable. Such a coating offers, moreover, the advantage of improved handling ability since the thus-treated propellant will no longer be dusty.

The mixtures to be utilized according to this invention of the tetrazole and/or its derivatives with the compounds from groups (A), (B) and/or (C) permit a finely graded staggering of the propellants with regard to their reaction rate and the thus-formed vapors and gases. This is a staggering needed for providing a maximum variety of uses for the propellant charges of this invention. Depending on the given structure of the generator housing of an airbag, for example, or of a belt tightening device, the propellant charges of this invention must be blended in a controlled fashion. Only in this way is it possible to attain an effect that is at an optimum in each case. The degree of efficiency of the propellant charges according to this invention is, after all, affected not only by the composition, but furthermore by the ignition, and also by the tamping resulting from the construction, and by the flow-off behavior of the thus-evolving vapors and gases. Evaluation of the degree of efficiency can be accomplished, for example, by determination of the gas pressure rise gradient of the respective

mixture in the given external environment in each case dependent on the structure, and the type of ignition selected. The thus-formed gas concentrations, specially those of the toxic gases, must not exceed specific maximum values. These values result from the MAK (maximale Arbeitsplatzkonzentration = maximum working site concentration) values (or TLV (threshold-limit value) values in the United States). Technical requirements are derived from these values in conjunction with the exposure times permitted and must be met by the respective propellant charges.

When determining these requirements, the differing passenger cells, for example, also enter into the consideration. In order to fulfill these demands, the respective propellant charge must be mixed in controlled fashion.

The values indicated in the tables were determined by igniting, in a cup in a loose bulk, 4 g of the respective propellant charge by means of a hot wire in a pressurized bomb tube having a volume of 25 ml. After ignition, a pressure-time curve was recorded. The thus-obtained values were evaluated as follows:

(1) Maximum pressure (bar): with the weighed-in amounts being the same, differences can be ascribed directly to the thus-produced gas volumes. These are determined by gas yield and thermal content of the reaction.

(2) Rise in gas pressure for the range of 40-60% of maximum pressure: In this region, the curve is no longer falsified by the ignition and/or the cooling-off characteristic of the vapors by



the inner surface of the bomb tube. The indicated times in  
millisconds (ms) represent the pressure rise and signify  
differing reaction rates. Such values are also obtained in the  
respective cases of application, for example in the diverse,  
5 structurally modified, gas generator housings. They permit a  
precise tuning of the propellant charges of this invention with a  
view toward the efficiency requirements. By the indication of  
the pressure rise times in a range of 40-60% of the maximum  
pressure, the reaction rate of the propellant charges of the  
10 invention is adequately characterized. The time periods for the  
reaction until the occurrence of maximum pressure serve as  
additional information.

The gas temperature can be kept low in a controlled fashion  
by addition of diammonium oxalate, oxalic acid diamide,  
15 dicyandiamide, or carbonates and, respectively, bicarbonates. If  
thermal stability is not a factor, and smoke formation upon  
addition of inorganic carbonates or bicarbonates is to be  
avoided, it is possible to use aminoguanidine bicarbonate as the  
organic bicarbonate. Other additives can be constituted by  
20 oxalic acid or urea, generally added in an amount of up to 5% by  
weight, based on the mixture.

Suitable reducing agents are metal powders of iron,  
magnesium, zirconium or titanium which, in contrast to the  
nonmetallic compound boron, have no strong effect on the  
25 deflagration velocity, but in case of the latter do have an  
influence on the heat production of the reaction and on the

reaction products. The proportion of the reducing agents can amount to up to 5% by weight.

Suitable catalysts for a mixture of such heterogeneous composition are compounds having an effect on the decomposition of peroxides, such as, for example, metals or their oxides, e.g. manganese dioxide. Additions of vanadium pentoxide or cerium dioxide lead to an increase in the reaction velocity. When adding up to 5% by weight of molybdenum(VI) oxide, the velocity is only slightly changed, and this also applies in the presence of cerium(III) nitrate hexahydrate. These additives are effective in amounts of up to a few percent by weight. Further catalysts are metal complexes, of which ferrocene can here be cited as an example; the addition of this compound with up to about 3% by weight markedly raises the reaction velocity.

The gas compositions described in accordance with this invention are manufactured by mixing the components according to conventional methods, optionally with the preparation of a harmless premix to which additional components are added. This mixture can already be utilized in powdered form. Unmixing on account of varying densities of the components can be counteracted by granulation of the mixture.

In the majority of applications by far, the mixture will be formed by press-molding or similar measures. To simplify this procedure, press-molding aids can be added to the mixture.

Suitable as such are graphite, molybdenum disulfide, "Teflon", talc, zinc stearate or boron nitride. These agents have an

effect even in minimum quantities and affect the properties and deflagration behavior either not at all or only to a minor extent.

It may be advantageous in some cases to influence the deflagration characteristic of the press-molded item by producing porosity. Such a method resides in admixing additives, such as salts, to the mixture prior to the actual shaping step; these additives can be removed again after molding by extraction with water or solvents. Another method resides in adding materials of low thermal resistance which decompose when the molded item is heated. The surface of the mixture can also be enlarged by adding to the mixture, prior to press-molding, hollow microspheres of glass or synthetic resins. The density of the press-molded article that can be attained thereby can deviate by approximately up to 20% from that of the untreated press-molded article; in this connection, this value is to be merely a roughly governing value and does not represent a limitation. This treatment leads to an extreme acceleration of the deflagration process.

A further treatment of the molded articles can consist in surface coating. In this case, besides achieving an effect on the deflagration characteristic, protection against environmental influences is obtained, in particular. Such a measure can also be suitable for increasing the strength of the molded item. In extreme cases, the use of suitable fibers for stabilization purposes would have to be additionally provided for. A side

effect of the coating procedure is a decreased abrasion of the items during transport stresses.

The thus-treated molded articles can be introduced in loose bulk or in oriented fashion into appropriate pressure-proof containers. They are ignited according to conventional methods with the aid of initiator charges or thermal charges wherein the thus-formed gases, optionally after flowing through a suitable filter, lead to inflation of the lifesaving system within fractions of a second.

The propellants of this invention are especially suited for so-called airbags, impact bags which are utilized in automotive vehicles or airplanes for occupants' protection. In case of vehicle impact, the airbag must fill up within a minimum time period with gas quantities of about 50 to 300 liters, depending on system and automobile size. The propellants of this invention are likewise suitable for use in belt-tightening devices.

Lifesaving systems containing the propellants of this invention likewise form the subject matter of the present invention.

#### Example 1

167 g of 5-aminotetrazole (5-ATZ) (produced from aminoguanidine sulfate, sodium nitrite and nitric acid) is recrystallized from about 600 ml of water under continuous agitation, dried at 110° C after filtration, ground up, and separated from coarse proportions with a 250  $\mu$ m screen (5-ATZ

specification: mp/decomp.:  $> 203^{\circ} \text{C}$ , average grain size  $80 \mu\text{m}$  and  $\text{H}_2\text{O}$  proportion  $< 0.05\%$ ). From  $\text{ZnSO}_4 \times 7 \text{H}_2\text{O}$  and hydrogen peroxide in aqueous ammonia,  $\text{ZnO}_2$  is produced, washed with dilute acetic acid as well as water, and dried at  $60^{\circ} \text{C}$  (specification of  $\text{ZnO}_2$ : 13.47% by weight of active oxygen, average grain size  $20.3 \mu\text{m}$ ).

5-ATZ and  $\text{ZnO}_2$ , as the components for nontoxic gas charges, are homogenized together in a weight ratio of 1 to 7 (corresponding to a molar ratio of about 1 : 5) in plastic containers in a tumbler mixer for 1-2 hours. Of this sample, 3.0 g is made to react in bulk in a 25 ml size stainless steel bomb tube by means of an electrically heatable Fe wire, and the pressure-time curve is recorded by means of a piezoelectric measuring device. After about 30 ms, a maximum gas pressure of about 200 bar is produced, due primarily to the formation of  $\text{CO}_2$ ,  $\text{N}_2$ ,  $\text{O}_2$  and  $\text{H}_2\text{O}$ . The reaction has a strongly exothermal character of about 471 cal/g (about 1970 J/g).  $\text{ZnO}$  remains as the reaction residue. The CO proportion corresponds to requirements. The deflagration point is at  $219^{\circ} \text{C}$ , the friction sensitivity is at 240 N, the impact sensitivity is at 20 J.

#### Examples 2-24

The use of 5-ATZ and  $\text{ZnO}_2$  as components in nontoxic gas charges corresponds to Example 1 with the use of further

additiv s. Examples 2-24 below describe the reaction of other  
mixtur s produced by conventional procedures. The results are  
compiled in Tabl s 1-4.

5 In the Tables, the first comment in the column labeled  
"Reaction Velocity" refers to Example No. 2, the second to  
Example No. 3 and so on.

TABLE 1

Example No.	Organic Additives (Moles)										Reaction Velocity (as Measured in Example 1)
	2	3	4	5	6	7	8	9	10		
5-Aminotetrazole	1	1	1	1	1	1	1	1	1	Rose as compared with 1	
Zinc peroxide	3	3	3	3	3	3	3	3	3	Rose as compared with 2	
Ammonium nitrate	2.5	2.5	2.5	2.5	2.5	2.5	2.5	2.5	2.5	Rose as compared with 1	
Aminoguanidine nitrate	1									Rose as compared with 2	
Urea		1								Dropped as compared with 2	
Oxalic acid dinitrate			1							Dropped as compared with 2	
Oxalic acid diamide				1						Dropped as compared with 2	
Biammonium oxalate monohydrate					1					Dropped as compared with 2	
Semicarbazide nitrate						1				Rose as compared with 2	
Aminoguanidine bicarbonate								1		Dropped as compared with 2	
Ferrocene									$1 \times 10^{-3}$	Rose as compared with 2	

TABLE 2

Example No.	Inorganic Additives (Moles)						Reaction Velocity (as Measured in Example 1)
	11	12	13	14	15	16	
5-Aminotetrazole	1	1	1	1	1	1	
Zinc peroxide	3	3	3	3	3	1.5	
Ammonium nitrate	1.5	2.5	2.5			2.5	
Iron nitrate nonahydrate	$1.6 \cdot 10^{-2}$						Dropped as compared with 2
Zinc carbonate	$2 \cdot 10^{-3}$						Dropped as compared with 2
Molybdenum(VI) oxide	$3 \cdot 10^{-3}$						Comparable to 2
Sodium nitrate				1			Dropped as compared with 2
Strontium nitrate					1		Dropped as compared with 2
Calcium peroxide						1.5	Comparable to 2



TABLE 3

Example No.	Catalytic Additives (Moles)							Reaction Velocity (as Measured in Example 1)
	17	18	19	20	21	22	23	
5-Aminotetrazole	1	1	1	1	1	1	1	1
Zinc peroxide	3	3	3	3	3	3	3	3
Ammonium nitrate	2.5	2.5	2.5	2.5	2.5	2.5	2.5	2.5
Vanadium(V) oxide	4·10 <sup>-5</sup>							Rose as compared with 2
Cerium(IV) oxide		5·10 <sup>-5</sup>						Rose as compared with 2
Cerium(IV) nitrate hexahydrate			4·10 <sup>-5</sup>					Comparable to 2
Manganese(IV) oxide			9·10 <sup>-5</sup>					Comparable to 2
Titanium					2·10 <sup>-3</sup>			Comparable to 2
Iron						1·10 <sup>-3</sup>		Comparable to 2
Magnesium							3·10 <sup>-3</sup>	Comparable to 2
Boron							7·10 <sup>-3</sup>	Rose as compared with 2

### Example 25

The gas charge mixtures described in Examples 1-24 can also be utilized in press-molded form. A mixture of 10 g of 5-ATZ ( $H_2O$  proportion < 0.1%, mp [decomposition] > 203° C, grain size 200-250  $\mu m$ ), 43.9 g of  $ZnO_2$  (12.85% by weight of active oxygen, grain size about 14  $\mu m$ ) and 23.5 g of  $NH_4NO_3$  (mp 167-169° C, grain size 250-315  $\mu m$ ), molar ratio 1 : 3 : 2.5, is blended in accordance with Example 1 and pressed into tablets (diameter = 6 mm, height = 2.77 mm, density = 2.18 g/ $CM^3$ , radial pressure force = 155.5  $\pm$  28.4 N) with a pressure of 4 tons. The deflagration behavior of the press-molded items, as tested in correspondence with Example 1, is slower than that of the bulk material and requires 0.1 g of B/ $KNO_3$  or Ti/ $ZnO_2$  as initiating mixture. The reaction velocity rises with a drop in the pressing force and drops with the size of the pressed material. The residue from the reactions remains extensively preserved in the form of the pressed items.

### Examples 26-32

As described in Example 1, further mixtures were prepared from gas-generating components and oxygen-yielding compounds, such as zinc peroxide, with an active oxygen proportion of 13.07% by weight and an average grain size of 11.8  $\mu m$  or, in case of sodium nitrate, with an average grain size of < 45  $\mu m$ .

Table 4 below contains additional data on the mixtures.

TABLE 4

Component	Melting Point (° C)	Molar Ratios or Mixture						
		26	27	28	29	30	31	32
5-aminotetrazole	206-208	1						
Potassium 5-amino- tetrazolate	269		1					
Bis(aminoguanidine)- azotetrazole	224-226			1				
Nitroguanidine	252				1			
Guanidine nitrate	210-214					1		
Semicarbazide nitrate	115-119						1	
1-Cyanoguanidine	203-210							1
Zinc peroxide	200 (decomp.)	2.48	2.43	10.64	1.42	1.42	0.71	4.25
Sodium nitrate	-	0.83	0.83	3.55	0.47	0.47	0.24	1.42

The components were homogenized in containers of plastic for a 1/2 hour with a tumbler-mixer, a 1/2 hour with a vibrator, and again for a 1/2 hour with a tumbler-mixer.

Of the thus-homogenized mixture, 4 g was introduced as described in Example 1 into a stainless steel pressure bomb tube and, under seal, made to react upon ignition with an incandescent wire. The following items were measured:

- the arising pressure (bar) up to the maximum value,
- the time (milliseconds, ms) up to maximum pressure value,
- the pressure increase gradient (dp/dt) between a pressure reaching 40-60% of the maximum pressure value.

The rise time served as the yardstick.

Table 5 below shows the values for the maximum pressure (bar) and the time in ms up to maximum pressure, ranging within the regions as described in Example 1 for a gas charge of 5-aminotetrazole and zinc peroxide. In addition, the time between 40 and 60% of maximum pressure was determined.

TABLE 5

Example No.	Max. Pressure (bar)	to Max. Pressure	Time (ms) to 40-60% of Max. Pressure
26	359	30	1.2
27	217	123	13.1
28	352	29	1.5
29	473	39	1.3
30	549	14	0.5
31	917	7	0.2
32	148	220	20.1

By adaption of the parameters and admixture of further components, it is possible to set the specifications required for the respectiv gas charge.

Another sample of the previously recited mixtures was studied with regard to physical and safety properties. The results are set forth in Table 6.

TABLE 6

Ex No.	Deflagra- tion Point	Friction Sensitivities	Impact Sensitivities	Heat of Explosion
	(°C) *	(N)	(J)	(J/g)
-----				
26	180	>360	7.5	2451
27	207	>360	10	2293
28	197	>360	4	2411
29	215	>360	20	2964
30	364	>360	15	2777
31	210	>360	2	3128
32	194	>360	30	2101

\*Determination took place according to the method of "Bundesanstalt fuer Materialpruefung" (BAM) [Federal Institute for Testing Materials] in Berlin.

The components are suitable for the production of gas charges on account of their miscibility, processability, press-molding ability for shaping, as well as compatibility with one another and with other additives, as well as due to their characteristic safety data.

Examples 33-44

As described in Examples 26-32, the mixtures of Examples 33-44 were produced from zinc peroxide (active oxygen proportion 12.0% by weight, average grain size  $4.8\ \mu\text{m}$ ), aminotetrazole (average grain size  $\leq 125\ \mu\text{m}$ ), sodium nitrate (grain size  $\leq 45\ \mu\text{m}$ ), and the listed components with a grain size of  $\leq 125\ \mu\text{m}$ .

The friction sensitivity, measured according to the method of BAM, was in all cases  $> 360\ \text{N}$ . The additionally listed components are disclosed in the literature.

Tables 7 and 8 below contain additional data on the mixtures.

TABLE 7

	Formulations	% by Wt.	Molar Pro- portions
Ex.No. 33	5-ATZ	29.8	1.4
	Zinc peroxide	23.4	0.75
	Sodium nitrate	46.8	2.2
Ex.No. 34	5-ATZ	19.5	1.0
	Dicyandiamidine nitrate	15.2	0.4
	Zinc peroxide	21.5	0.75
	Sodium nitrate	43.8	2.24
Ex.No. 35	5-ATZ	18.1	1.0
	Dicyandiamidine sulfate	12.8	0.2
	zinc peroxide	19.8	0.75
	Sodium nitrate	49.3	2.73
Ex.No. 36	5-ATZ	19.5	1.0
	1-Cyanoguanidine	7.7	0.4
	Zinc peroxide	21.4	0.75
	Sodium nitrate	51.4	2.64
Ex.No. 37	5-ATZ	16.9	1.0
	Melamine	10.0	0.4
	Zinc peroxide	18.6	0.75
	Sodium nitrate	54.5	3.22
Ex.No. 38	5-ATZ	20.2	1.0
	Azodicarboxylic acid diamide	11.0	0.4
	Zinc peroxide	22.2	0.75
	Sodium nitrate	46.6	2.31
Ex.No. 39	5-ATZ	19.6	1.0
	Cyanuric acid	11.9	0.4
	Zinc peroxide	21.4	0.75
	Sodium nitrate	47.1	2.41
Ex.No. 40	5-ATZ	22.5	1.0
	Urea	6.4	0.4
	Zinc peroxide	24.7	0.75
	Sodium nitrate	46.4	2.06

TABLE 7 (Continued)

	Formulations	% by Wt.	Molar Pro- portions
Ex.No. 41	5-ATZ	20.2	1.0
	Biuret	9.8	0.4
	Zinc peroxide	22.2	0.75
	Sodium nitrate	47.8	2.37
Ex.No. 42	5-ATZ	21.0	1.0
	Aminoguanidine nitrate	13.5	0.4
	Zinc peroxide	23.0	0.75
	Sodium nitrate	42.5	2.03
Ex.No. 43	5-ATZ	20.5	1.0
	Sodium dicyanamide	8.6	0.4
	Zinc peroxide	22.5	0.75
	Sodium nitrate	48.4	2.37
Ex.No. 44	5-ATZ	23.9	1.0
	Sodium cyanate	7.3	0.4
	Zinc peroxide	26.2	0.75
	Sodium nitrate	42.6	1.79



T A B L E 8

Mixture Components in Molar Proportions	Example No.	Deflagration Point (° C)	Impact Sensitivity (J)	Heat of Explosion (J/g)	Maximum Pressure (bar)	Gas Pressure Ris for Range of 40-60% p <sub>max</sub> (ms)	Sodium Nitrate	5-Amino- tetrazole	Zinc Peroxide
Dicyandiamidine nitrate	33	>400	5	3258	683	0.38	2.2	1.4	0.75
Dicyandiamidine sulfate	34	>400	6	3142	761	0.40	2.24	1.0	0.75
l-Cyanoguanidine	35	395	7.5	2883	656	1.04	2.73	1.0	0.75
Melamine	36	367	10	3038	661	0.36	2.64	1.0	0.75
Azodicarboxylic acid diamide	37	>400	10	3187	652	1.16	3.22	1.0	0.75
Cyanuric acid	38	>400	6	3191	706	0.36	2.31	1.0	0.75
Urea	39	>400	7.5	2732	582	0.80	2.41	1.0	0.75
Biuret	40	>400	10	3053	654	0.40	2.06	1.0	0.75
Aminoguanidine nitrate	41	363	7.5	2982	663	0.56	2.37	1.0	0.75
Sodium dicyanamide	42	256	7.5	3190	693	0.30	2.03	1.0	0.75
Sodium cyanate	43	356	7.5	3226	486	0.36	2.37	1.0	0.75
	44	349	10	3005	458	0.34	1.79	1.0	0.75